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(54) Title: ELECTROLYTE SYSTEM AND ENERGY STORAGE DEVICE USING SAME

(57) Abstract: A secondary cell employs a nonaqueous electrolyte solution including a nonaqueous solvent and a salt, and a flame retardant material that is liquid at room temperature and pressure and substantially immiscible in the nonaqueous electrolyte solution. The nonaqueous electrolyte solution is formed by dissolving a salt, preferably an alkali metal salt, in a nonaqueous solvent. The nonaqueous solvent preferably includes a cyclic carbonate and/or a linear carbonate. The cyclic carbonate preferably contains an alkylene group with 2 to 5 carbon atoms, and the linear carbonate preferably contains a hydrocarbon group with 1 to 5 carbon atoms. Preferred salts include LiPF<sub>6</sub> and LiBF<sub>4</sub> at a concentration between about 0.1 and 3.0 moles/liter. The flame retardant material is preferably a halogen-containing compound in an amount by weight of nonaqueous solvent in a range of about 1 to about 99 wt%, and preferred halogen-containing compounds contain perfluoralkyl or perfluorether groups.

## **ELECTROLYTE SYSTEM AND ENERGY STORAGE DEVICE USING SAME**

### **REFERENCE TO PRIOR-FILED APPLICATIONS**

This application claims priority to U.S. Patent Application Serial Number 10/034,316,  
5 filed December 28, 2001.

### **BACKGROUND**

Energy storage devices such as batteries that utilize the oxidation and reduction  
reactions of an alkali metal are known. Such "lithium ion cells" include secondary cells  
which use a carbon material that can be doped and undoped with lithium ions as a negative  
10 pole and which use a complex oxide of lithium and a second metal as a positive pole.  
Lithium ion cells are compact and lightweight, and have a high energy density. Accordingly,  
the use of lithium ion cells as secondary cells in portable electronic devices has expanded  
rapidly. Concurrently, there has been an escalating demand for improved performance, e. g.,  
an increase in the energy density and an increase in the discharge current, etc., in lithium ion  
15 cells in order to achieve a further improvement in the function of such portable electronic  
devices.

The decreased size of such energy storage devices has resulted in the presence of  
highly energetic active substances in a small, confined volume. As a result, large amounts of  
energy can be released when electrodes short-circuit or otherwise fail as a result of, for  
20 example, piercing and compression that may cause a battery to ignite and catch fire. As the  
cell capacity has increased, there has been a strong demand for an improvement in battery  
safety.

Previous attempts to increase safety have been directed to changing electrode  
construction or changing the active substance. Other studies have focused on additives to the  
25 electrolyte solution that increase the safety of the batteries. Thus, for example phosphorus

and fluorine compounds such as triphenyl phosphate and fluoro-ethers have been dissolved in the electrolyte solution to improve battery safety. However, these compounds may be subject to oxidation-reduction by the electrodes, or may react with the electrodes so that the capacity is lowered. Although safety is improved when the amounts of additives are increased, battery  
5 performance deteriorates. Accordingly, it has been difficult to realize increased safety without causing a deterioration of conventional battery characteristics.

There is thus a continuing need for improved cell capacity, charge-discharge rate, and charge-discharge cycle.

There is also a continuing need for an improved battery with an improved safety  
10 profile that does not deteriorate over time.

#### SUMMARY

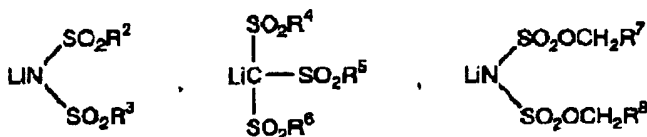
One object of the present invention is to provide an electrolyte system for an energy storage device that is extremely safe and has superior charge-discharge characteristics.

Another object of the present invention is to provide an electrolyte system of high  
15 conductivity and low viscosity for improving the discharge capacity of a secondary cell battery.

Another object of the present invention is to provide an electrolyte system that is chemically and electrically stable with respect to the positive pole and negative pole of the secondary cell battery.

20 In accordance with the present invention, an electrolyte system comprises a nonaqueous electrolyte solution including a nonaqueous solvent and a salt, and a flame retardant material that is a liquid at room temperature and pressure and that is substantially immiscible in the nonaqueous electrolyte solution.

The nonaqueous electrolyte solution is formed by dissolving a salt, preferably an alkali metal salt, in a nonaqueous solvent. The nonaqueous solvent is a polar aprotic organic solvent which readily dissolves alkali metal salts, and which is resistant to being electrolyzed by oxidation-reduction at the battery poles. The nonaqueous solvent preferably includes a  
 5 cyclic carbonate and/or a linear carbonate, wherein the cyclic carbonate preferably contains an alkylene group with 2 to 5 carbon atoms, and the linear carbonate preferably contains a hydrocarbon group with 1 to 5 carbon atoms. Preferred electrolytes include  $\text{LiPF}_6$ ,  $\text{LiBF}_4$ ,  $\text{LiOSO}_2\text{R}^1$ ,



10 (in the above formulae,  $\text{R}^1$  through  $\text{R}^8$  indicate perfluoroalkyl, alkyl or aryl groups with 1 to 6 carbon atoms, which may be the same or different groups). The concentration of the electrolytes in the nonaqueous electrolyte solution is preferably between about 0.1 to 3.0 moles/liter, and more preferably between about 0.5 to 2.0 moles/liter.

The flame retardant material is a liquid at room temperature and pressure and is  
 15 substantially immiscible in the nonaqueous electrolyte solution. Preferably, the flame retardant material is a halogen-containing compound. Preferred halogen-containing compounds contain perfluoroalkyl groups or perfluorether groups. The halogen-containing compound is present in an amount by weight of nonaqueous solvent in a range of from about 1 to about 99 wt%, preferably from about 1 to about 70 wt%, even more preferably from  
 20 about 10 to about 60 wt%, and even more preferably from about 20 to about 40 wt%.

An energy storage device in accordance with the present invention comprises the disclosed electrolyte system, an electrode assembly including a first electrode member, a

second electrode member, a separator member physically and electrically separating the first electrode member from the second electrode member but capable of allowing ionic conductivity between the first electrode member and the second electrode member through the nonaqueous electrolyte solution, and a casing enclosing the electrode assembly and the electrolyte system. In a energy storage device, such as a battery, the first electrode member is a negative electrode containing a material selected from the group consisting of lithium metal, a lithium alloy, a carbon material that can be doped and undoped with lithium ions, a metal oxide that can be doped and undoped with lithium ions, and silicon that can be doped and undoped with lithium ions; the second electrode member is a positive electrode containing a material selected from the group consisting of complex oxide of lithium and a transition metal, and a complex oxide of lithium, transition metal and a non-transition metal; and the separator member is a resin containing a polymer.

A method of making a energy storage device in accordance with the present invention comprises providing an electrode assembly including a first electrode member, a second electrode member and a separator member physically and electrically separating the first electrode member from the second electrode member but capable of allowing ionic conductivity between the first electrode member and the second electrode member, placing the assembly in a casing, and filling the casing with the electrolyte system of the present invention by first, filling the casing at least partially with the nonaqueous electrolyte solution, waiting a period of time sufficient for the nonaqueous electrolyte solution to penetrate one or more pores of the electrode assembly, and then adding the flame retardant material to the casing.

In practice, the two phase electrolyte system of the present invention realizes significant advantages over the prior art. Because the halogen-containing compound is not miscible in the nonaqueous electrolyte solution, there is little ingress of the halogen-

containing compound into the regions within the casing occupied by the nonaqueous electrolyte solution, i.e. in the casing regions separating the positive and negative poles and defined by the separator. As such, there is little interference by the halogen-containing compound in electrochemical reactions occurring at the poles and in the nonaqueous electrolyte solution. Further, the secondary cell is superior in terms of initial capacity and cycle characteristics. Moreover, the halogen-containing compound of the present invention interferes and/or inhibits combustion reactions that may result upon piercing or compression of the battery. Accordingly, the energy storage device of the present invention exhibits enhanced safety over prior art secondary cells.

## 10 DETAILED DESCRIPTION

The following is a list of references that may be relevant to the present application.

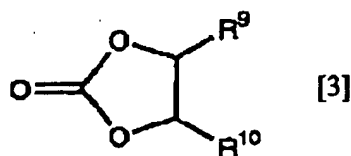
US	5,916,708		06/29/1999	Besenhard et al.
EP	0 807 986	A1	11/19/1997	Hoechst Aktiengesellschaft
EP	1 205 997	A1	05/15/2002	Bridgestone Corporation
EP	1 253 662	A1	10/30/2002	Bridgestone Corporation
JP	10050343	A2	02/20/1998	Hoechst AG
JP	11317232	A2	11/16/1999	Mitsubishi Chemical Corp.
JP	2000106209	A2	04/11/2000	Mitsui Chemicals Inc.
JP	2000164249	A2	06/16/2000	Mitsui Chemicals Inc.
JP	2001060464	A2	03/06/2001	Japan Storage Battery Co. Ltd.
JP	2001143749	A2	05/25/2001	Nippon Zeon Co. Ltd.
JP	2002042891	A2	02/08/2002	Toshiba Battery Co. Ltd.
JP	2002190316	A2	07/05/2002	Mitsubishi Chemicals Corp.

An electrolyte system for an energy storage device includes a nonaqueous electrolyte solution including a nonaqueous solvent and a salt, and a flame retardant material that is a liquid at room temperature and pressure and that is substantially immiscible in the

nonaqueous electrolyte solution. The nonaqueous electrolyte solution permits the exchange of metal ions between the positive pole and negative pole of the battery.

The nonaqueous electrolyte solution of the present invention is formed by dissolving a salt, preferably an alkali metal salt, in a nonaqueous solvent. The nonaqueous solvent of the present invention is preferably a polar aprotic organic solvent which readily dissolves alkali metal salts, and which is resistant to being electrolyzed by oxidation-reduction at the poles.

The nonaqueous solvent of the present invention preferably includes one or more cyclic carbonates of the following general formula:

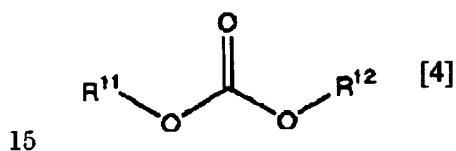


where  $R^9$  and  $R^{10}$  may be the same or different, and may be hydrogen, a linear alkyl group, a branched alkyl group, a cyclic alkyl group, or a halogen-substituted alkyl group in which some or all of the hydrogen atoms are substituted by at least one halogen selected from a set consisting of chlorine and bromine. When linear alkyl groups are selected, linear alkyl groups with 1 to 4 carbon atoms such as methyl groups, ethyl groups, propyl groups and butyl groups, etc., are preferred. When branched alkyl groups are selected, branched alkyl groups with 3 to 10 carbon atoms, and especially 3 to 6 carbon atoms, such as isopropyl groups, isobutyl groups, sec-butyl groups and tert-butyl groups, etc., are preferred. When cyclic alkyl groups are selected, cyclic alkyl groups with 5 to 10 carbon atoms such as cyclopentyl groups, cyclohexyl groups and 1-methylcyclohexyl groups, etc., are preferred.

The nonaqueous solvent may include cyclic carbonates having 5-member ring compounds as shown in above formula and 6-member ring compounds. Examples of suitable cyclic carbonates include ethylene carbonate, propylene carbonate, 1,2-butylene carbonate, 2,3-butylene carbonate, 1,3-propylene carbonate, 1,3-butylene carbonate, 2,4-pentylene carbonate, 1,3-pentylene carbonate, and vinylene carbonate, etc. Furthermore, halogen-substituted cyclic carbonates in which some or all of the hydrogen atoms in the methyl groups of the aforementioned propylene carbonates, etc., are substituted by at least one halogen selected from a set consisting of fluorine, chlorine, and bromine may also be used.

Preferably, the cyclic carbonate contains an alkylene group with 2 to 5 carbon atoms. More preferably, ethylene carbonate and propylene carbonate are used. Two or more such cyclic carbonates may be used in combination.

The nonaqueous solvent of the present invention may additionally include one or more linear carbonates shown in the following general formula:



where R<sup>11</sup> and R<sup>12</sup> may be the same or different, and may be branched or cyclic alkyl groups, or halogen-substituted alkyl groups in which some or all of the hydrogen atoms are substituted by at least one halogen selected from a set consisting of fluorine, chlorine and bromine. When a linear alkyl group is selected, linear alkyl groups with 1 to 4 carbon atoms such as methyl groups, ethyl groups, propyl groups, and butyl groups, etc., are preferred. When branched alkyl groups are selected, branched alkyl groups with 3 to 10 carbon atoms, such as isopropyl groups, isobutyl groups, sec-butyl groups, and tert-butyl groups, etc., are



preferred. When cyclic alkyl groups are selected, cyclic alkyl groups with 5 to 10 carbon atoms such as cyclopentyl groups, cyclohexyl groups, and 1-methylcyclohexyl groups, etc., are preferred.

Examples of suitable linear carbonates include dimethyl carbonate, diethyl carbonate,  
5 di-n-propyl carbonate, dibutyl carbonate, diisopropyl carbonate, and methylethyl carbonate, etc. Among such linear carbonates, linear carbonates that contain a hydrocarbon group with 1 to 5 carbon atoms are desirable in the present invention, and dimethyl carbonate, methylethyl carbonate, and diethyl carbonate are especially desirable.

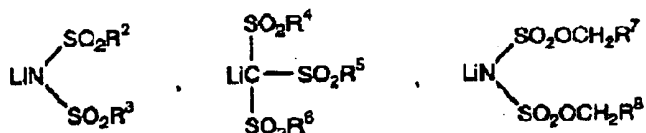
The nonaqueous solvent of the present invention may also include one or more  
10 solvents selected from the group consisting of cyclic esters such as  $\gamma$ -butyrolactone,  $\gamma$ -valerolactone, 3-methyl- $\gamma$ -butyrolactone and 2-methyl- $\gamma$ -butyrolactone, etc., linear esters such as methyl formate, ethyl formate, methyl acetate, ethyl acetate, propyl acetate, methyl propionate, methyl butyrate and methyl valerate, etc., cyclic ethers such as 1,4-dioxane, 1,3-dioxolane, tetrahydrofuran, 2-methyltetrahydrofuran, 3-methyl-1,3-dioxolane and 2-methyl-  
15 1,3-dioxolane, etc., linear ethers such as 1,2-dimethoxyethane, 1,2-diethoxyethane, diethyl ether, dimethyl ether, methylethyl ether and dipropyl ether, etc., sulfur-containing compounds such as sulfolane and dimethyl sulfate, etc., and phosphorus-containing compounds such as trimethylphosphoric acid and triethylphosphoric acid, etc.

The nonaqueous solvent of the present invention may be one solvent used singly or  
20 may be a mixture consisting of two or more solvents.

One or more salts are dissolved in the nonaqueous solvent in order to prepare the nonaqueous electrolyte solution of the present invention. Preferably, the concentration of the salts in the nonaqueous electrolyte solution is between about 0.1 to 3.0 moles/liter, and more preferably between about 0.5 to 2.0 moles/liter. The salt used in the present invention may be

any salt known to those of ordinary skill in the art to be used in nonaqueous electrolyte solutions of energy storage devices such as secondary cells. The salt may be a single salt or a mixture of salts. Suitable examples of salts include lithium salts such as  $\text{LiPF}_6$ ,  $\text{LiBF}_4$ ,  $\text{LiClO}_4$ ,  $\text{LiAsF}_6$ ,  $\text{LiAlCl}_4$ ,  $\text{Li}_2\text{SiF}_6$ ,  $\text{LiOSO}_2\text{R}^1$ ,

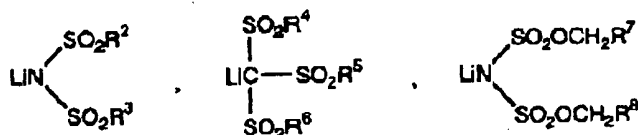
5



where  $\text{R}^1$  through  $\text{R}^8$  may be the same or different, and are perfluoroalkyl, alkyl, or aryl groups with 1 to 6 carbon atoms, and alkali metal salts in which the lithium in the above-mentioned salts is substituted by alkali metals. These compounds may be used singly or in

10 mixtures consisting of two or more compounds.

Preferred electrolytes are  $\text{LiPF}_6$ ,  $\text{LiBF}_4$ ,  $\text{LiOSO}_2\text{R}^1$ ,



15 The electrolyte system of the present invention further includes a flame retardant material. A suitable flame retardant material is liquid at room temperature and pressure, is substantially anhydrous and is substantially immiscible with the electrolyte solution. Substantially immiscible means that when the nonaqueous electrolyte solution and the halogen-containing compound are mixed, a meniscus between the two liquids is formed and

the concentration of the halogen-containing compound in the nonaqueous electrolyte solution is about 10 wt % or less and preferably, 1000 ppm or less. More preferably, the flame retardant material is completely immiscible in the nonaqueous electrolyte solution. Suitable flame retardant materials are stable with respect to electro-chemical reactions on the electrodes, and are chemically inert with respect to the electrolyte and the nonaqueous solvent in the battery environment. Suitable flame retardant materials are further characterized by having very low vapor pressure at room temperature.

Preferably, the flame retardant material is a halogen containing compound. Suitable halogen-containing compounds include branched or unbranched alkyls, cyclic alkyls, ethers, amino alkyls, aliphatic heterocyclic compounds in which one or more of the hydrogen atoms are substituted with a halogen atom selected from the group consisting of fluorine, chlorine, and bromine. In general, increasing efficacy is obtained with increases in the number of halogen atoms. Preferably, the halogen-containing compounds include perfluoroamine and perfluoroether compounds. Preferred halogen-containing compounds include perfluoro-1,3-dimethylcyclohexane and FC-70, sold under the trade name Fluorinert, having a molecular formula  $C_{15}F_{33}N$ , and commercially available from 3M Co. of St. Paul, Minnesota.

Preferably, the halogen-containing compound is present in an amount by weight of nonaqueous solvent in a range of from about 1 to about 99% by weight of nonaqueous solvent, preferably in a range of from about 1 to about 70% by weight of nonaqueous solvent, more preferably from about 10 to about 60% by weight of nonaqueous solvent, and even more preferably from about 20 to about 40% by weight of nonaqueous solvent.

In the electrolyte systems of the present invention, the use of the halogen-containing compound and a cyclic carbonate and/or linear carbonate is especially preferred. In cases where a halogen-containing compound and a cyclic carbonate and/or linear carbonate is used,

the halogen-containing compound is preferably present in an amount by weight of nonaqueous solvent ranging from about 1 to about 99 wt %, preferably from about 1 to about 70 wt %, more preferably from about 10 to about 60 wt %, and even more preferably from about 20 to about 40 wt %. Furthermore, it is desirable that the aforementioned cyclic carbonate and/or linear carbonate be contained in the nonaqueous solvent in an amount from about 1 to 99 wt %, preferably from about 30 to 99 wt %, and even more preferably from about 40 to 90 wt %.

The electrolyte system of the present invention is used in connection with an electrode assembly and a casing to produce an energy storage device such as a battery. An energy storage device according to the present invention comprises an electrolyte system in accordance with the present invention; an electrode assembly including a first electrode member, a second electrode member, and a separator member physically and electrically separating the first electrode member from the second electrode member but capable of allowing ionic conductivity between the first electrode member and the second electrode member through the nonaqueous electrolyte solution; and a casing enclosing the electrode assembly and the electrolyte system.

Preferably the energy storage device in accordance with one embodiment of the present invention is a battery in which the first electrode member is a negative electrode containing a material selected from the group consisting of lithium metal, a lithium alloy, a carbon material that can be doped and undoped with lithium ions, a metal oxide such as tin oxide, titanium oxide, or silicon oxide that can be doped and undoped with lithium ions, and silicon that can be doped and undoped with lithium ions, or a mixture thereof. The preferred materials for the negative pole are carbon, silicon, silicon oxide, and mixtures thereof. Preferred carbon for the negative pole includes graphite, amorphous carbon, carbon fibers, and meso-carbon micro-beads.

In this embodiment, the second electrode member is a positive electrode containing a material selected from the group consisting of complex oxide of lithium and a transition metal, and a complex oxide of lithium, transition metal and a non-transition metal. Suitable materials include complex oxides of lithium and transition metals, such as  $\text{LiCoO}_2$ ,  $\text{LiMnO}_2$ ,  
5  $\text{LiMn}_2\text{O}_4$ ,  $\text{LiNiO}_2$ ,  $\text{LiNi}_x\text{Co}_{(1-x)}\text{O}_2$ ,  $\text{LiNiMnCoO}_2$ ,  $\text{LiFePO}_4$ ,  $\text{LiNiCoAlO}_2$ , vanadium oxide and lithium vanadium oxide.

The separator member may be a porous polymer or ceramic material. Preferably, the separator is a resin containing a polymer, generally a microporous polymer film. Suitable examples include polyesters, polyacetyl, and polyethylene.

10 Energy storage devices may be prepared using the electrolyte system of the present invention as follows. A method of making an energy storage device in accordance with the present invention comprises providing an electrode assembly including a first electrode member, a second electrode member, and a separator member physically and electrically separating the first electrode member from the second electrode member but capable of  
15 allowing ionic conductivity contact between the first electrode member and the second electrode member, placing the assembly in a casing, and filling the casing with the electrolyte system that is herein described, filling the casing at least partially with the nonaqueous electrolyte solution, waiting a period of time sufficient for the nonaqueous electrolyte solution to penetrate one or more pores of the electrode assembly, and then adding the flame  
20 retardant material to the casing.

The energy storage devices of the present invention may take any suitable shape or configuration, including a battery of suitable shape and configuration, including a coin type battery, a prismatic battery, a stacked battery, cylindrical battery, etc.

Within the battery casing, the nonaqueous electrolyte solution selected in accordance with the present invention and the halogen-containing compound selected in accordance with the present invention form two separate liquid phases. Since the nonaqueous electrolyte solution and the halogen-containing compound are substantially immiscible, one liquid phase within the battery casing is mainly nonaqueous electrolyte solution and the other liquid phase in the battery casing is mainly halogen-containing compound. In accordance with the present invention, the nonaqueous electrolyte solution phase occupies the region within the casing between the positive pole and the negative pole defined by the separator. As a result, the halogen-containing compound is not in the region between the positive pole and the negative pole defined by the separator and does not interfere with the oxidation-reduction processes at the poles and in the nonaqueous electrolyte solution. However, the halogen-containing compound tends to inhibit the occurrence of highly energetic events such as combustion processes and therefore increases battery safety. Applicant believes the halogen-containing compound may interfere with the production of highly energetic and highly reactive carbon radicals produced in combustion processes and/or interfere with chain reaction processes by production of fluoride radicals.

The following examples are illustrative of certain embodiments of the invention and are not to be construed as limiting the invention thereto.

#### Example 1

A positive pole prepared from  $\text{LiCoO}_2$  and poly(vinylidene fluoride) (PVDF) and a negative pole prepared from graphite and PVDF were fastened together with a separator (manufactured by Tonen) between them, thus producing an electrode assembly. The assembly was then introduced into a battery casing.

A nonaqueous electrolyte solution was prepared by mixing ethylene carbonate (EC), diethyl carbonate (DEC), and  $\text{LiPF}_6$  in a weight ratio of EC to DEC to  $\text{LiPF}_6$  of 21.84 wt%:

37.59 wt%: 10.57 wt%. The nonaqueous electrolyte solution was then introduced into the battery casing. After waiting 1 hour, the flame retardant material, a halogen-containing compound, perfluoro-1,3-dimethylcyclohexane was then introduced into the battery casing at a 30% weight ratio of halogen-containing compound to electrolyte solution.

- 5           The battery was sealed and charged at 0.1 C to a charge of 4.3 V by constant-current constant-voltage charging. The battery was then discharged at a constant current and the discharge capacity and safety level measured. The results are shown in Table 1.

#### Example 2

- 10           A nonaqueous electrolyte solution and battery were prepared in the same manner as in Example 1, except that FC-70 was used as the halogen-containing compound. The battery discharge capacity and safety level were measured and the results are shown in Table 1.

#### Comparative Example 1

- 15           A nonaqueous electrolyte solution was prepared in the same manner as in Example 1, except that a halogen-containing compound was not added to the electrolyte solution. The battery discharge capacity and safety level were measured and the results are shown in Table 1.

- The safety of the energy storage devices was assessed using an "Overcharge Test." The above examples were overcharged at a 4.4-4.5 V. The temperature on the outside of the battery casing was measured. Overcharging was continued until the battery became mechanically detached or until the temperature of the battery casing peaked and began to decrease. Safety was assessed according to a 4 level designation:
- 20

Level 1:       Maximum Casing Temperature less than 80°C

Level 2:       Maximum Casing Temperature between 80°C and 120°C

Level 3:       Maximum Casing Temperature over 120°C

- 25       Level 4:       Mechanical Detachment

As indicated by the results tabulated in Table 1, addition of the halogen-containing compound to the nonaqueous electrolyte solution significantly improved the safety level of the battery (from "4" to either "2" or "1") without adversely affecting the discharge capacity of the battery to any significant degree.

- 5 The principle preferred embodiments and modes of operation of the present invention have been described. The invention described herein, however, is not intended to be construed as limited to the particular forms disclosed, since they are regarded as illustrative rather than restrictive. Variations and changes may be made by those skilled in the art without departing from the spirit of the invention.

10

Table 1

Additive name	Discharge capacity(%)	Safety level	EC ratio(wt%)	DEC ratio(wt%)	LiPF6 ratio(wt%)	Additive ratio(%)
none	100.0	4	31.2	53.7	15.1	0
perfluoro-1,3-dimethylcyclohexane	96.3	1	21.84	37.59	10.57	30
	93.7	2	21.84	37.59	10.57	30
FC-70	99.5	2	21.84	37.59	10.57	30
	100.5	2	21.84	37.59	10.57	30

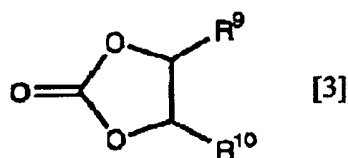


## CLAIMS

I claim:

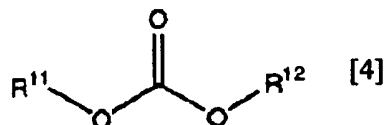
1. An electrolyte system comprising:  
  
a nonaqueous electrolyte solution including a nonaqueous solvent and a salt; and  
  
5 a flame retardant material that is a liquid at room temperature and pressure and substantially immiscible in the nonaqueous electrolyte solution.
2. The electrolyte system according to claim 1, wherein the salt is present in the nonaqueous electrolyte solution in a concentration ranging from about 0.1 to about 3.0 moles/liter.
- 10 3. The electrolyte system according to claim 1, wherein the salt is present in the nonaqueous electrolyte solution in a concentration ranging from about 0.5 to about 2.0 moles/liter.
4. The electrolyte system according to claim 1, wherein the flame retardant material is a halogen-containing compound.
- 15 5. The electrolyte system according to claim 4, wherein the halogen-containing compound is present in an amount ranging from about 1 to about 99% by weight of the nonaqueous solvent.
6. The electrolyte system according to claim 4, wherein the halogen-containing compound is present in an amount ranging from about 1 to about 70% by weight of the  
  
20 nonaqueous solvent.
7. The electrolyte system according to claim 4, wherein the halogen-containing compound is present in an amount ranging from about 10 to about 60% by weight of the nonaqueous solvent.

8. The electrolyte system according to claim 4, wherein the halogen-containing compound is present in an amount ranging from about 20 to about 40% by weight of the nonaqueous solvent.
9. The electrolyte system of claim 1, wherein the nonaqueous solvent includes at least one carbonate selected from the group consisting of cyclic carbonates, linear carbonates and mixtures thereof.
10. The electrolyte system of claim 9, wherein the cyclic carbonate contains an alkylene group with 2 to 5 carbon atoms and the linear carbonate contains a hydrocarbon group with 1 to 5 carbon atoms.
11. The electrolyte system according to claim 1, wherein the nonaqueous solvent includes at least one cyclic carbonate represented by the formula



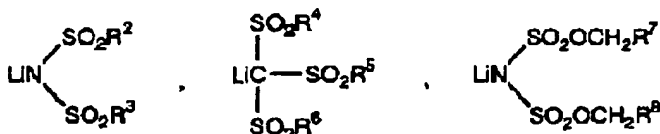
- wherein  $R^9$  and  $R^{10}$  are independently selected from the group consisting of: hydrogen, linear alkyl groups, branched alkyl groups, cyclic alkyl groups, and halogen-substituted alkyl groups in which at least one hydrogen is substituted by chlorine or bromine.

12. The electrolyte system according to claim 1, wherein the nonaqueous solvent includes at least one linear carbonate represented by the formula



wherein  $R^{11}$  and  $R^{12}$  are independently selected from the group consisting of linear alkyl groups, branched alkyl groups, cyclic alkyl groups, and halogen-substituted alkyl groups in which at least one hydrogen atom is substituted by fluorine, chlorine, or bromine.

13. The electrolyte system according to claim 1, wherein the salt is selected from the  
5 group consisting of  $LiPF_6$ ,  $LiBF_4$ ,  $LiOSO_2R^1$ ,



wherein  $R^1$  through  $R^8$  are independently selected from the group consisting of: perfluoroalkyls, alkyls, and aryls with 1 to 6 carbon atoms.

- 10 14. The electrolyte system according to claim 4, wherein the halogen-containing compound contains at least one member selected from the group consisting of: branched or unbranched alkyl, cyclic alkyl, ether, aminoalkyl, and aliphatic heterocyclic compound groups in which one or more hydrogen atoms are substituted by a halogen selected from the group consisting of: fluorine, chlorine, and bromine.
- 15 15. The electrolyte system according to claim 4, wherein the halogen-containing compound contains at least one member selected from the group consisting of: perfluoroalkyl groups, perfluoroaminoalkyl groups, perfluoroether groups, and mixtures thereof.
16. The electrolyte system according to claim 4, wherein the halogen-containing compound is perfluoro-1,3-dimethylcyclohexane.
- 20 17. A method of making an energy storage device comprising:  
providing an electrode assembly including:

a first electrode member;

a second electrode member; and

a separator member physically and electrically separating the first electrode member from the second electrode member but capable of allowing ionic conductivity

5 between the first electrode member and the second electrode member;

placing the assembly in a casing; and

filling the casing with the electrolyte system according to any of claims 1 to 15 by first, filling the casing at least partially with the nonaqueous electrolyte solution, waiting a period of time sufficient for the nonaqueous electrolyte solution to penetrate one or more pores of the electrode assembly, and then adding the flame retardant material to the casing.

18. The method of making an energy storage device of claim 17, further comprising the steps of charging energy storage device after adding the nonaqueous electrolyte solution.

19. A method of making an energy storage device according to claim 17, wherein the first electrode member is a negative electrode, the second electrode member is a positive  
15 electrode, and the separator member is a resin containing a polymer.

20. An energy storage device comprising:

the electrolyte system according to any of claims 1 to 15;

an electrode assembly, the electrode assembly including:

a first electrode member;

20 a second electrode member; and

a separator member physically and electrically separating the first electrode member from the second electrode member but capable of allowing ionic conductivity

between the first electrode member and the second electrode member through the nonaqueous electrolyte solution; and

a casing enclosing the electrode assembly and the electrolyte system.

21. The energy storage device according to claim 20, wherein:

5 the first electrode member is a negative electrode containing a material selected from the group consisting of lithium metal, a lithium alloy, a carbon material that can be doped and undoped with lithium ions, a metal oxide that can be doped and undoped with lithium ions, and silicon that can be doped and undoped with lithium ions;

10 the second electrode member is a positive electrode containing a material selected from the group consisting of: complex oxide of lithium and a transition metal, and a complex oxide of lithium, transition metal, and a non-transition metal; and

the separator member is a resin containing a polymer.

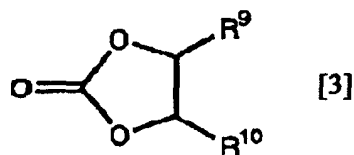
22. The energy storage device according to claim 21, wherein said negative electrode contains a material selected from the group consisting of carbon, silicon, silicon oxide, and  
15 mixtures thereof.

23. The energy storage device according to claim 21, wherein the energy storage device is a lithium battery.

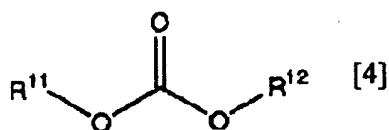
## AMENDED CLAIMS

[received by the International Bureau on 01 May 2003 (01.05.03);  
original claims 10, 17, 18 and 20 amended,  
remaining claims unchanged (4 pages)]

8. The electrolyte system according to claim 4, wherein the halogen-containing compound is present in an amount ranging from about 20 to about 40% by weight of the nonaqueous solvent.
9. The electrolyte system of claim 1, wherein the nonaqueous solvent includes at least one carbonate selected from the group consisting of cyclic carbonates, linear carbonates and mixtures thereof.
10. The electrolyte system of claim 9, wherein the at least one carbonate comprises a carbonate selected from the group consisting of: a cyclic carbonate containing an alkylene group with 2 to 5 carbon atoms and a linear carbonate containing a hydrocarbon group with 1 to 5 carbon atoms.
11. The electrolyte system according to claim 1, wherein the nonaqueous solvent includes at least one cyclic carbonate represented by the formula



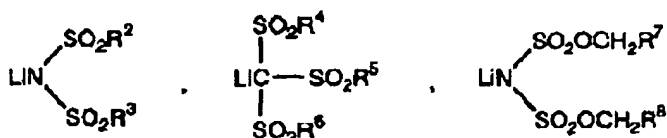
- 15 wherein  $R^9$  and  $R^{10}$  are independently selected from the group consisting of: hydrogen, linear alkyl groups, branched alkyl groups, cyclic alkyl groups, and halogen-substituted alkyl groups in which at least one hydrogen is substituted by chlorine or bromine.
12. The electrolyte system according to claim 1, wherein the nonaqueous solvent includes at least one linear carbonate represented by the formula



17 (replacement 14 April 2003)

wherein  $R^{11}$  and  $R^{12}$  are independently selected from the group consisting of linear alkyl groups, branched alkyl groups, cyclic alkyl groups, and halogen-substituted alkyl groups in which at least one hydrogen atom is substituted by fluorine, chlorine, or bromine.

13. The electrolyte system according to claim 1, wherein the salt is selected from the group consisting of  $LiPF_6$ ,  $LiBF_4$ ,  $LiOSO_2R^1$ ,



wherein  $R^1$  through  $R^8$  are independently selected from the group consisting of: perfluoroalkyls, alkyls, and aryls with 1 to 6 carbon atoms.

14. The electrolyte system according to claim 4, wherein the halogen-containing compound contains at least one member selected from the group consisting of: branched or unbranched alkyl, cyclic alkyl, ether, aminoalkyl, and aliphatic heterocyclic compound groups in which one or more hydrogen atoms are substituted by a halogen selected from the group consisting of: fluorine, chlorine, and bromine.
15. The electrolyte system according to claim 4, wherein the halogen-containing compound contains at least one member selected from the group consisting of: perfluoroalkyl groups, perfluoroaminoalkyl groups, perfluoroether groups, and mixtures thereof.
16. The electrolyte system according to claim 4, wherein the halogen-containing compound is perfluoro-1,3-dimethylcyclohexane.
17. A method of making an energy storage device comprising:
- providing an electrode assembly including:

18 (replacement 14 April 2003)

a first electrode member;

a second electrode member; and

a separator member physically and electrically separating the first electrode member from the second electrode member but capable of allowing ionic conductivity

5 between the first electrode member and the second electrode member;

placing the assembly in a casing; and

filling the casing with the electrolyte system according to any of claims 1 – 15, or 24 by first, filling the casing at least partially with the nonaqueous electrolyte solution, waiting a period of time sufficient for the nonaqueous electrolyte solution to penetrate one or more  
10 pores of the electrode assembly, and then adding the flame retardant material to the casing.

18. The method of making an energy storage device of claim 17, further comprising the steps of charging the energy storage device after filling the casing at least partially with the nonaqueous electrolyte solution and before adding the flame retardant material to the casing.

19. A method of making an energy storage device according to claim 17, wherein the first  
15 electrode member is a negative electrode, the second electrode member is a positive electrode, and the separator member is a resin containing a polymer.

20. An energy storage device comprising:

the electrolyte system according to any of claims 1 – 15, or 24;

an electrode assembly, the electrode assembly including:

20 a first electrode member;

a second electrode member; and

19 (replacement 14 April 2003)



a separator member physically and electrically separating the first electrode member from the second electrode member but capable of allowing ionic conductivity between the first electrode member and the second electrode member through the nonaqueous electrolyte solution; and

5 a casing enclosing the electrode assembly and the electrolyte system.

21. The energy storage device according to claim 20, wherein:

the first electrode member is a negative electrode containing a material selected from the group consisting of lithium metal, a lithium alloy, a carbon material that can be doped and undoped with lithium ions, a metal oxide that can be doped and undoped with lithium  
10 ions, and silicon that can be doped and undoped with lithium ions;

the second electrode member is a positive electrode containing a material selected from the group consisting of: complex oxide of lithium and a transition metal, and a complex oxide of lithium, transition metal, and a non-transition metal; and

the separator member is a resin containing a polymer.

15 22. The energy storage device according to claim 21, wherein said negative electrode contains a material selected from the group consisting of carbon, silicon, silicon oxide, and mixtures thereof.

23. The energy storage device according to claim 21, wherein the energy storage device is a lithium battery.

20 24. The electrolyte system according to claim 1, wherein the flame retardant material is  $C_{15}F_{33}N$ .

20 (replacement 14 April 2003)

# INTERNATIONAL SEARCH REPORT

International application No.

PCT/US02/39451

## A. CLASSIFICATION OF SUBJECT MATTER

IPC(7) : H01M 10/40

US CL : 429/306

According to International Patent Classification (IPC) or to both national classification and IPC

## B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

U.S. : 429/306, 324, 326, 330, 332, 338, 342; 252/62.2

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)  
Please See Continuation Sheet

## C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X --- Y	US 5,830,600 A (NARANG et al) 03 November 1998 (03.11.1998), column 10, lines 17-27; column 11, lines 39-44; column 12, lines 42-51; column 20, lines 50-60; column 22, lines 25-35.	1-16, 20-23 ----- 17-19
X --- Y	JP 10-012272 A (MITSUBISHI CHEM CORP.) 16 January 1998 (16.01.1998), paragraphs 4-6, 13, 14, 17, 18, 20, 21, 24-31, and 42-44 of JPO Machine translation.	1-16, 20-23 ----- 17-19

☐ Further documents are listed in the continuation of Box C.

☐ See patent family annex.

\* Special categories of cited documents:

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"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

"Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art

"&" document member of the same patent family

Date of the actual completion of the international search

23 February 2003 (23.02.2003)

Date of mailing of the international search report

20 MAR 2003

Name and mailing address of the ISA/US

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**INTERNATIONAL SEARCH REPORT**

PCT/US02/39451

**Continuation of B. FIELDS SEARCHED Item 3:**

**EAST (USPAT, US-PGPUB, EPO, JPO, DERWENT)**

search terms: electrolyte, flame, fire, retardant, inflammable, incombustible, miscibility, immiscibility

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